

Bushong: Radiologic Science for Technologists, 10th Edition

Chapter 02: The Structure of Matter

Test Bank

MULTIPLE CHOICE

1. The term “atom” was first used by the _____.
 - a. Ethiopians
 - b. British
 - c. Greeks
 - d. Romans

ANS: C

The term “atom” was first used by the Greeks

DIF: Moderate REF: page 27 OBJ: Relate the history of the term “atom.”

2. The first person to describe an element as being composed of identical atoms was _____.
 - a. J. J. Thomson
 - b. John Dalton
 - c. Dmitri Mendeleev
 - d. Niels Bohr

ANS: B

The first person to describe an element as being composed of identical atoms was John Dalton.

DIF: Moderate REF: page 27

OBJ: Name the first person to describe an element as being composed of identical atoms.

3. The smallest particle that has all the properties of an element is a(n) _____.
 - a. neutron
 - b. proton
 - c. electron
 - d. atom

ANS: D

The smallest particle that has all the properties of an element is an atom.

DIF: Moderate REF: page 28 OBJ: Define the atom.

4. The periodic table of the elements was developed by _____ in the late 19th century.
 - a. Bohr

- b. Rutherford
- c. Mendeleev
- d. Roentgen

ANS: C

The Periodic Table was developed by Mendeleev.

DIF: Moderate REF: page 28

OBJ: Name the person who developed the periodic table of the elements.

5. Rutherford's experiments in 1911 showed that the atom was composed of _____.
a. electrons with well-defined orbits
b. a nucleus with an electron cloud
c. electrified plum pudding
d. a ball of hooks and eyes

ANS: C

Rutherford's experiments in 1911 showed that the atom was composed of a nucleus with an electron cloud.

DIF: Moderate REF: page 29

OBJ: Relate the history of the Rutherford model of the atom.

6. A positively charged nucleus surrounded by negatively charged electrons in well-defined orbits is the _____ model of the atom.
a. Bohr
b. Thomson
c. Rutherford
d. Dalton

ANS: A

A positively charged nucleus surrounded by negatively charged electrons in well-defined orbits is the Bohr model of the atom.

DIF: Moderate REF: page 29

OBJ: Identify the structure of the Bohr model of the atom.

7. What are the fundamental particles of an atom?
a. quark, positron, negatron
b. nucleon, electron, proton
c. proton, neutron, quark
d. proton, electron, neutron

ANS: D

The fundamental particles of an atom are the proton, electron, and neutron.

DIF: Easy REF: page 29 OBJ: Identify the fundamental particles of an atom.

8. The chemical element is determined by the number of _____ in the atom.
- protons
 - electrons
 - neutrons
 - nucleons

ANS: A

The chemical element is determined by the number of protons in the atom.

DIF: Moderate REF: page 30 OBJ: Describe how a chemical element is determined.

9. An atom in a normal state has an electrical charge of _____.
- one
 - zero
 - positive
 - negative

ANS: B

An atom in a normal state has an electrical charge of zero.

DIF: Moderate REF: page 31
OBJ: Describe the electrical charge of an atom in a normal state.

10. The binding energies, or energy levels, of electrons are represented by their _____.
- atomic numbers
 - atomic mass units
 - shells
 - isotopes

ANS: C

The binding energies, or energy levels, of electrons are represented by their shells.

DIF: Moderate REF: page 33
OBJ: Describe binding energies or energy levels of electrons.

11. When an atom has the same number of protons as another, but a different number of neutrons, it is called an _____.
- isomer
 - isobar
 - isotone
 - isotope

ANS: D

When an atom has the same number of protons as another, but a different number of neutrons, it is called an isotope.

DIF: Difficult REF: page 34 OBJ: Describe an isotope.

12. When atoms of various elements combine, they form _____.
a. isotopes
b. compounds
c. molecules
d. ions

ANS: C

When atoms of various elements combine, they form molecules.

DIF: Moderate REF: page 36 OBJ: Describe a molecule.

13. An atom that loses or gains one or more electrons is a(n) _____.
a. ion
b. molecule
c. isotope
d. isomer

ANS: A

An atom that loses or gains one or more electrons is an ion.

DIF: Moderate REF: page 31 OBJ: Define an ion.

14. The maximum number of electrons that can exist in an electron shell is calculated with the formula _____.
a. $2n$
b. $2n^2$
c. $2/n$
d. $2/n^2$

ANS: B

The number of electrons in an electron shell is calculated with the formula $2n^2$.

DIF: Difficult REF: page 32

OBJ: Identify the formula for the maximum number of electrons that can exist in an electron shell.

15. A neutral atom has the same number of _____ and electrons.
a. quarks
b. neutrinos
c. neutrons

d. protons

ANS: D

A neutral atom has the same number of protons and electrons.

DIF: Easy REF: page 31

OBJ: Identify the formula for the maximum number of electrons that can exist in an electron shell.

16. The innermost electron shell is symbolized by the letter ____.

- a. J
- b. K
- c. L
- d. M

ANS: B

The innermost electron shell is symbolized by the letter K.

DIF: Moderate REF: page 31

OBJ: Recognize the symbol for the innermost electron shell.

17. The shell number of an atom is called the ____.

- a. alpha particle
- b. chemical element
- c. principal quantum number
- d. half-life number

ANS: C

The shell number of an atom is called the principal quantum number.

DIF: Moderate REF: page 32 OBJ: Define the shell number of an atom.

18. The atomic number of an element is symbolized by the letter ____.

- a. A
- b. X
- c. Z
- d. n

ANS: C

The atomic number of an element is symbolized by the letter Z.

DIF: Moderate REF: page 34

OBJ: Identify symbol for the atomic number of an element.

19. Aluminum has an atomic number of 13. How many protons does it have?

- a. 13

- b. 26
- c. 27
- d. none of the above

ANS: A

The atomic number equals the number of protons in an atom.

DIF: Moderate REF: page 34

OBJ: Identify the number of protons on an atom based on its atomic number.

20. Two identical atoms which exist at different energy states are called ____.
- a. isotopes
 - b. isomers
 - c. isotones
 - d. isobars

ANS: B

Two identical atoms which exist at different energy states are called isomers.

DIF: Moderate REF: page 36 OBJ: Define an isomer.

21. The atomic number of molybdenum is 42 and the atomic mass number is 98. How many neutrons does it have?
- a. 42
 - b. 98
 - c. 21
 - d. 56

ANS: D

The number of neutrons is equal to $A - Z$.

DIF: Difficult REF: page 36

OBJ: Identify the number of neutrons in an atom based on its atomic number and atomic mass number.

22. A chemical compound is any quantity of ____.
- a. one type of atom
 - b. one type of molecule
 - c. two types of molecules
 - d. two or more types of atoms

ANS: B

A chemical compound is any quantity of one type of molecule.

DIF: Difficult REF: page 36 OBJ: Describe a compound.

23. During beta emission, an atom releases _____.
a. electrons
b. positrons
c. protons
d. neutrons

ANS: A

During beta emission, an atom releases electrons.

DIF: Moderate REF: page 37 OBJ: Describe beta emission.

24. The only difference between x-rays and gamma rays is their _____.
a. energy
b. size
c. origin
d. name

ANS: C

The only difference between x-rays and gamma rays is their origin.

DIF: Moderate REF: page 42

OBJ: Explain the difference between x-rays and gamma rays.

25. The _____ is the least penetrating form of ionizing radiation.
a. beta particle
b. x-ray
c. gamma ray
d. alpha particle

ANS: D

The alpha particle is the least penetrating form of ionizing radiation.

DIF: Moderate REF: page 41

OBJ: Name the least penetrating form of ionizing radiation.